

Give <sup>n</sup>conc., write ratio

$$.163 \% (m/m) \text{ KCl} \rightarrow \frac{.163 \text{ g KCl}}{100 \text{ g soln}}$$

$$7.0 \% (v/v) \text{ alc} \rightarrow \frac{7.0 \text{ mL alc}}{100 \text{ mL soln}}$$

$$56.8 \% (m/v) \text{ NaCl} \rightarrow \frac{56.8 \text{ g NaCl}}{100 \text{ mL soln}}$$

$$.525 \text{ M HCl} \rightarrow \frac{0.525 \text{ mole HCl}}{1 \text{ L soln}}$$

$$.64 \text{ m CaBr}_2 \rightarrow \frac{0.64 \text{ mole CaBr}_2}{1 \text{ kg H}_2\text{O}}$$

NOT CONCENTRATIONS

Given molar mass, write ratio

molar mass of HCl is 36.5 g/mole

$$\text{ratio: } \frac{36.5 \text{ g HCl}}{1 \text{ mole HCl}}$$

Given density, write ratio

density of Au is 19 g/cm<sup>3</sup> (or 19 g/mL)

$$\text{ratio: } \frac{19 \text{ g Au}}{1 \text{ cm}^3 \text{ Au}}$$

density of a soln is .975 g/mL

$$\text{ratio: } \frac{.975 \text{ g soln}}{1 \text{ mL soln}}$$