

EQUATIONS WORKSHEET

Name _____
Last First

For each of the following

1. Indicate in the blank to the left whether the type of reaction is (A) combination (B) decomposition (C) single replacement (D) double replacement (E) combustion
2. Predict the products and write the complete balanced equation for the reaction, giving the physical state { (s), (l), (g), (aq) } for each reactant and product.

_____ 1. Aqueous solutions of calcium chloride and silver nitrate are mixed

_____ 2. Sulfur dioxide is bubbled through liquid water.

_____ 3. Zinc is added to an aqueous solution of cupric nitrate.

_____ 4. Aqueous solutions of oxalic acid and sodium hydroxide are mixed.

_____ 5. Cupric hydroxide is heated.

_____ 6. Aqueous solutions of ferric chloride and potassium phosphate are mixed.

_____ 7. Gaseous butane, C_4H_{10} , reacts in air.

_____ 8. Sodium chloride is heated.

_____ 9. Magnesium is added to an aqueous solution hydrochloric acid.

_____ 10. Chlorine is mixed with aqueous barium bromide.

_____ 11. Magnesium carbonate is heated.

_____ 12. Decomposition of hydrogen peroxide.

_____ 13. Potassium oxide and carbon dioxide are mixed

_____ 14. Sodium is added to liquid water.

_____ 15. Sodium oxide is added to liquid water.

_____ 16. Sodium nitrate is heated.

_____ 17. Calcium oxide is added to liquid water.

_____ 18. Aqueous solutions of potassium iodide and lead nitrate are mixed.

_____ 19. Calcium hydroxide is added to an aqueous solution of sulfuric acid.

_____ 20. ~~Diphosphorus pentoxide and water forming phosphoric acid.~~

_____ 21. Calcium sulfite is heated.

_____ 22. Magnesium oxide and sulfur dioxide are mixed.

_____ 23. Calcium is added to an aqueous solution of sulfuric acid.

_____ 24. Magnesium oxide is added to liquid water.

_____ 25. Sodium peroxide is heated.

_____ 26. Potassium chlorate is heated.
