

WORKSHEET 6

Chemistry 110

Name _____

(last)

(first)

Due date: _____

A. Write formulas for the following compounds.

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|------------------------------------|--|-------------------------------|---|
| 1. calcium iodide | <u>CaI_2</u> | 11. lithium hydroxide | <u>LiOH</u> |
| 2. zinc hydroxide | <u>$\text{Zn}(\text{OH})_2$</u> | 12. dibromine pentasulfide | <u>Br_2S_5</u> |
| 3. trisulfur heptanitride | <u>S_3N_7</u> | 13. zinc nitrite | <u>$\text{Zn}(\text{NO}_2)_2$</u> |
| 4. stannous borate | <u>$\text{Sn}_3(\text{BO}_3)_2$</u> | 14. silver acetate | <u>$\text{AgC}_2\text{H}_3\text{O}_2$</u> |
| 5. oxalic acid | <u>$\text{H}_2\text{C}_2\text{O}_4$</u> | 15. ferric chromate | <u>$\text{Fe}_2(\text{CrO}_4)_3$</u> |
| 6. aluminum bromite | <u>$\text{Al}(\text{BrO}_2)_3$</u> | 16. sulfurous acid | <u>H_2SO_3</u> |
| 7. cobaltic phosphite | <u>CoPO_3</u> | 17. ammonium oxalate | <u>$(\text{NH}_4)_2\text{C}_2\text{O}_4$</u> |
| 8. cuprous cyanide | <u>CuCN</u> | 18. carbon tetrachloride | <u>CCl_4</u> |
| 9. mercury (I) nitrate | <u>$\text{Hg}_2(\text{NO}_3)_2$</u> | 19. nickel (III) permanganate | <u>$\text{Ni}(\text{MnO}_4)_3$</u> |
| 10. potassium dihydrogen phosphate | <u>KH_2PO_4</u> | 20. perchloric acid | <u>HClO_4</u> |

B. Write names for the following compounds: *SPELLING COUNTS!*

- | | |
|---|--------------------------------|
| 1. N_2H_4 | <u>dinitrogen tetrahydride</u> |
| 2. $\text{HIO}_2(\text{aq})$ | <u>iodous acid</u> |
| 3. $\text{HF}(\text{aq})$ | <u>hydrofluoric acid</u> |
| 4. CdCO_3 | <u>cadmium carbonate</u> |
| 5. SnSO_4 | <u>tin (II) sulfate</u> |
| 6. OBr_2 | <u>oxygen dibromide</u> |
| 7. $\text{Mg}(\text{ClO}_2)_2$ | <u>magnesium chlorite</u> |
| 8. Na_2O_2 | <u>sodium peroxide</u> |
| 9. $\text{H}_2\text{Cr}_2\text{O}_7(\text{aq})$ | <u>dichromic acid</u> |
| 10. IF_7 | <u>iodine heptafluoride</u> |